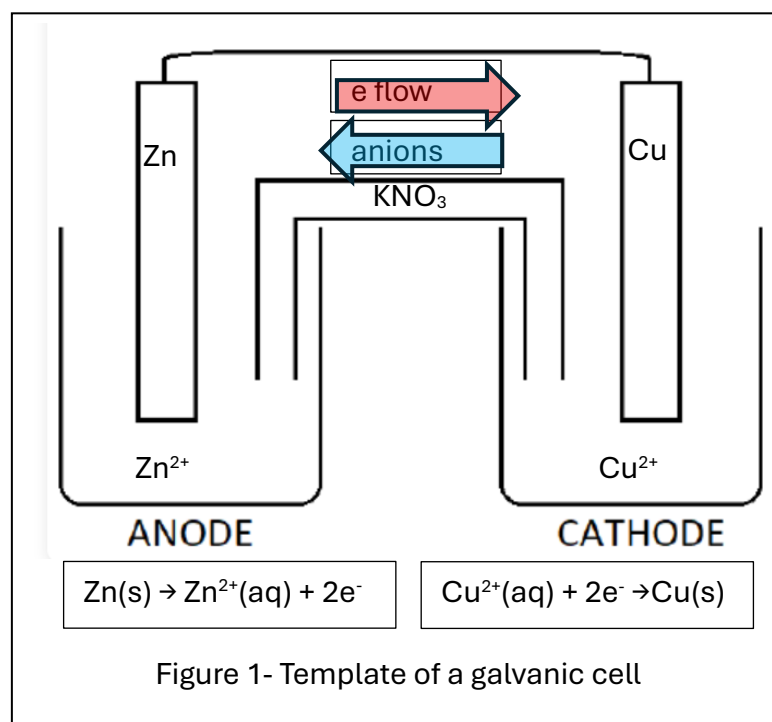


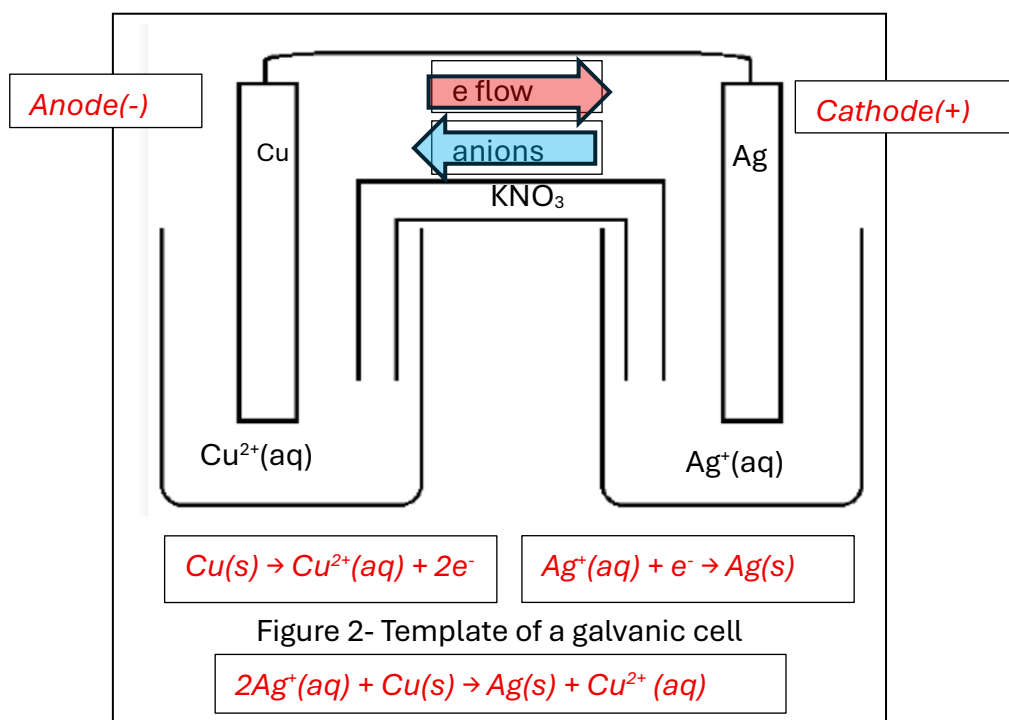
1. Consider the galvanic cell shown in fig. 1.



- Place the following half cells in the appropriate place in the template.
 $\text{Zn}^{2+}(\text{aq}) / \text{Zn}(\text{s})$ and $\text{Cu}^{2+}(\text{aq}) / \text{Cu}(\text{s})$
- Write the oxidation and reduction half equations in the appropriate boxes in the diagram in fig. 1.
- Indicate on the diagram in fig. 1 the direction of electron flow.
- Indicate in the box in fig 1 the direction of anion flow.
- Using item 2 of the 2026 VCAA Chemistry Data booklet find the EMF of the cell in fig. 1.

$$+0.34\text{V} - -0.76\text{V} = 1.10\text{V}$$

2. Consider the galvanic cell shown in fig. 2.



- Label the anode and cathode in the boxes provided in fig. 2
- Write the oxidation and reduction half equations in the appropriate boxes in the diagram in fig. 2.
- Indicate on the diagram in fig. 2 the direction of electron flow.
- Indicate in the box in fig. 2 the direction of anion flow.
- Using item 2 of the 2026 VCAA Chemistry Data booklet find the EMF of the cell.

$$+0.80V - +0.34V = +0.46 V$$

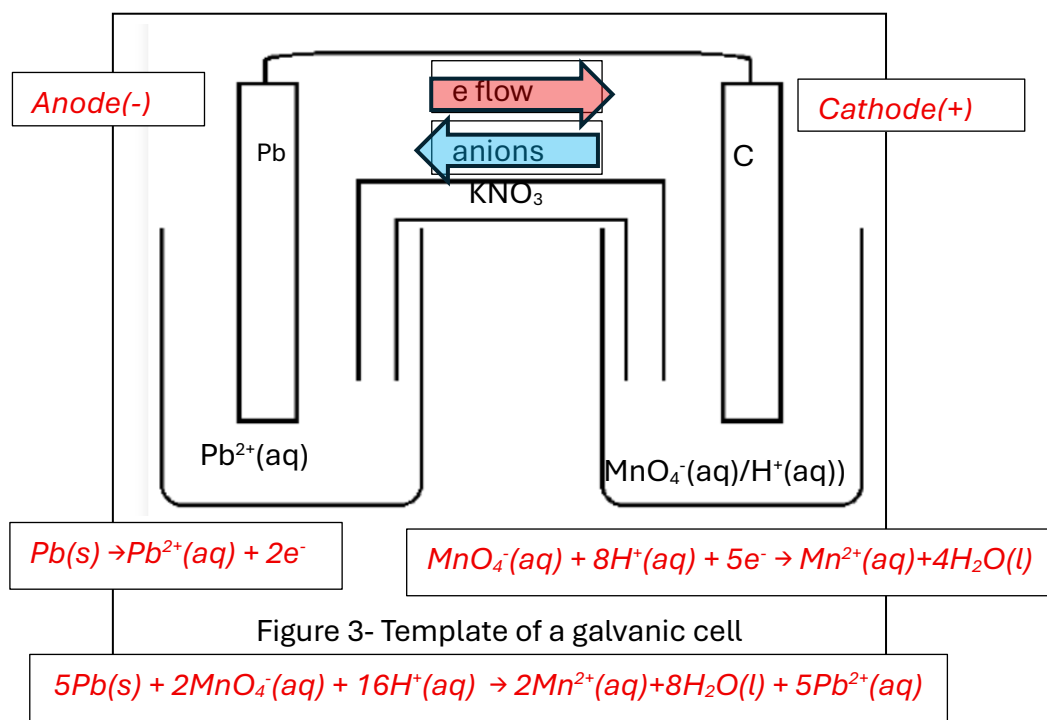
f. Write the :

- oxidation reaction _____

- reduction reaction _____

Overall redox equation _____

3. Consider the galvanic cell shown in fig. 3.



- Label the anode and cathode in the boxes provided in fig. 3
- Write the oxidation and reduction half equations in the appropriate boxes in the diagram in fig. 3.
- Indicate on the diagram in fig. 3 the direction of electron flow.
- Indicate in the box in fig. 3 the direction of anion flow.
- Using item 2 of the 2026 VCAA Chemistry Data booklet, find the EMF of the cell.

$$+1.51V - -0.13V = +1.64V$$

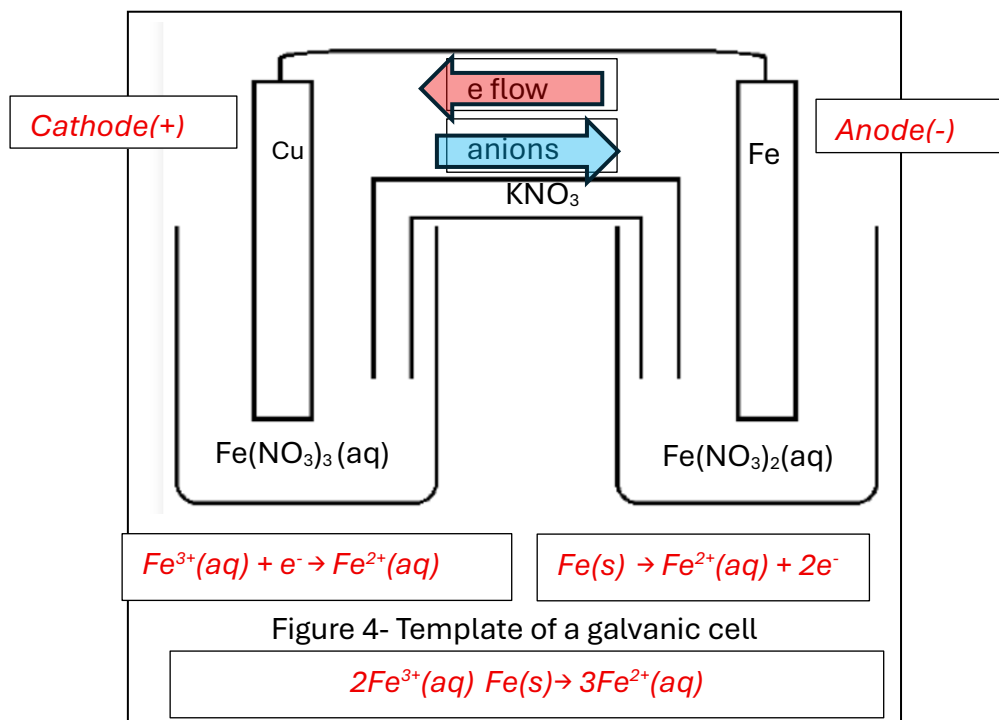
f. Write the :

- oxidation reaction _____

- reduction reaction _____

Overall redox equation _____

4. Consider the galvanic cell shown in fig. 4.



- Label the anode and cathode in the boxes provided in fig. 4
- Write the oxidation and reduction half equations in the appropriate boxes in the diagram in fig. 4.
- Indicate on the diagram in fig. 4 the direction of electron flow.
- Indicate in the box in fig. 4 the direction of anion flow.
- Using item 2 of the 2026 VCAA Chemistry Data booklet find the EMF of the cell.

$$+0.77\text{V} - -0.44\text{V} = +1.21\text{V}$$

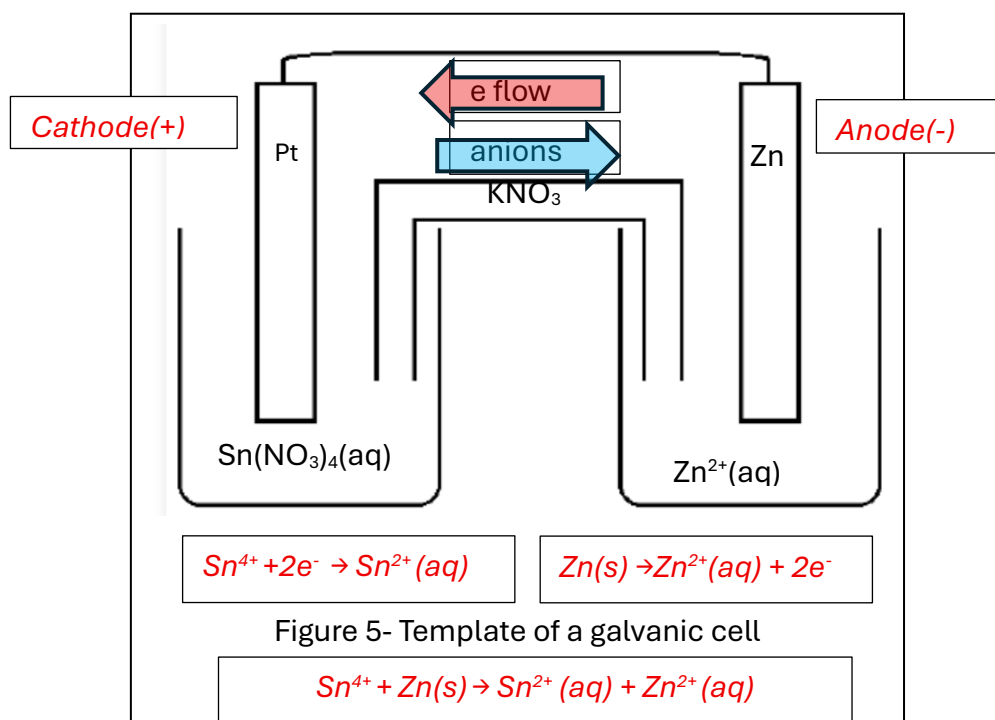
f. Write the :

- oxidation reaction _____

- reduction reaction _____

Overall redox equation _____

5. Consider the galvanic cell shown in fig. 5.



- Label the anode and cathode in the boxes provided in fig. 5
- Write the oxidation and reduction half equations in the appropriate boxes in the diagram in fig. 5.
- Indicate on the diagram in fig. 5 the direction of electron flow.
- Indicate in the box in fig. 5 the direction of anion flow.
- Using item 2 of the 2026 VCAA Chemistry Data booklet, find the EMF of the cell.

$$+0.15V - -0.76V = +0.91V$$

e. Write the :

- oxidation reaction _____

- reduction reaction _____

Overall redox equation _____